Iraqi Journal of Scientific and Industrial Research
Volume 1, Issue (4), October-December 2022, pp. 122-128

Biological Activity of Ca and Zn Complexes with Valine and 2,2′-bipyridine

Athraa M. JASEM

Department of Physics, College of Education, Al-Iraqia University, Baghdad, IRAQ

Abstract
Zinc(II) and Calcium(II) ions reacted with 2,2′-bipyridine as a primary ligand and L-valine as a secondary ligand to create two new mixed ligand complexes of transition metals. X-ray crystallography, ultraviolet visible spectra, and Fourier-transform infrared spectroscopy were used to analyze the ligands and their metal complexes. Zn(Val)2(bipy.) and Ca(Val)2(bipy.) were the equations used to characterize the mixed ligand complexes. The crystal structure symmetry of both complexes is indicated by the pseudo-octahedral structure of the metal complexes and their similar structural patterns. On the other hand, the yield of Ca complex was 81% of the basic materials, while the yield of Zn complex was 50% of them. Ca chelate's ability to control Phytophthora was demonstrated via bioactivity

Keywords: Metal complexes; 2,2′-bipyridine; Zinc ions; Calcium ions; Ligands

Received: 2 August 2021; **Revised:** 3 October 2021; **Accepted:** 10 October 2021; **Published:** 1 January 2022

1. Introduction

L-Valine is one of the 20 proteinogenic amino acids with the chemical formula HOCCH(NH)CH(CH) [1]. It is classified as a non-polar branched-chain amino acid and is widely present in human food sources such as cottage cheese, fish, poultry, peanuts, sesame seeds, and lentils [2]. However, it is seldom present at a ratio exceeding 10%. It can be obtained from alanine *via* adding two methyls (CH_3) groups to the α-carbon atom [3]. The properties and structure of valine are shown in Fig. (1) and table (1), respectively.

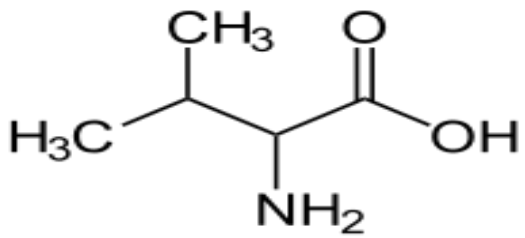


Fig. (1) Structure of valine

With regards to the structure of ligand complexes, the insertion of a second competing ligand such as 1,10-phenanthroline, 2,2-bipyridine diminishes the dimensions of the metal complex structure. The 2,2′-bipyridine ligand (Fig. 2) has been widely used as a metal chelating ligand especially when combined with other transition metals owing to its following properties (1) strong redox stability and ease of functionalization, and (2) neutrality that can constitute charged complexes with metal cations and form symmetrical and asymmetrical isomers [4]. The properties of 2,2′-bipyridine are shown in table (2).

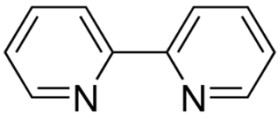


Fig. (2) Structure of 2,2′-bipyridine

Another part of complexes (Ca chelate or Zn chelate) represents transition metal ions

such as Fe, Co, Ni, Cu, Zn, Ca and Cd that play a vital role in living systems, such as in enzymes or in carriers in a macrocyclic ligand field environment. Bio-coordination chemistry studies into the role of transition metal ions in living systems are now of great interest to chemists [5]. Bio-coordination chemistry can improve our understanding about living systems and the use of these metal ions to create and/or prepare different metal complexes can have a range of applications in our daily life. For many years, the use of bio-coordination chemistry to produce fungicides was surprisingly neglected and currently is still in its early stages. The specific objective of this study was to synthesize new ligand complexes that can potentially use as fungicides against *Phytophthora* species and to investigate of their physical and chemical structural characteristics [6-9].

The study of ligand complexes has become increasingly important with respect to their biological activity toward pathogens [10,11]. In general, mixed ligand complexes have been extensively used as antimicrobial [12-14] and anticancer agents [15]. Recently, the importance of bioinorganic chemistry has grown globally, especially when European Union countries picked "Bio-coordination Chemistry" as one of the seven priority research fields in 1991 [16].

The product of potassium phosphite (K_2HPO_3) is used widely against *Phytophthora* species to protect natural ecosystems, urban trees and orchards [17-19]. Phosphite is translocated in both the xylem and the phloem through association with photo-assimilates in a source-sink relationship given that it is trapped inside the phloem [20-22]. Mechanisms of action of phosphite include inciting the plant defence responses and/or act on the pathogen directly by causing inhibition or death [23-26]. *Phytophthora* species cause a variety of diseases in numerous plant species in urban ecosystems [27-30]. In some cities, *Phytophthora* species are commonly associated with declining urban trees [31].

Chemical treatments are considered effective means for control of *Phytophthora* species that attack plants in natural or agricultural environments. Phosphite is a chemical widely used to manage the spread and impact of diseases caused by oomycete plant pathogens, especially *Phytophthora* species [32]. Unfortunately, excessive phosphite concentrations can result in phytotoxicity in plants. Also, some *Phytophthora* species have tolerance to phosphite.

2. Materials and Methods

2.1 Preparation of Ca complex

The Ca complex was made using a common method of making mixed ligands metal complexes [33,34]. Briefly, a solution of 2,2′-bipyridine (0.156 g, 1 m.mole) in aqueous ethanol (1:1:5 ml) and solution of L-Valine (0,234, 2 m.mole) in aqueous ethanol (1:1:5 ml) containing sodium hydroxide (0.08, 2 mmol) were added simultaneously to a solution of $CaCl_2.6H_2O$ (1 m.mole) in aqueous ethanol (1:1:10 ml) in the stoichiometric ratio [2Val:Ca:bipy]. The solution was stirred constantly at room temperature for 4 hours then allowed to stand overnight. The crystallized product was filtrated off and washed with aqueous ethanol.

2.2 Preparation of Zn complex

The Zn complex was made using the same method of making Ca complex [35,36]. Briefly, a solution of 2,2′-bipyridine (0.156 g, 1 m.mole) in aqueous ethanol (1:1:5 ml) and solution of L- Valine (0,234, 2 m.mole) in aqueous ethanol (1:1:5 ml) containing sodium hydroxide (0.08, 2mmol) were added simultaneously to solution of $ZnCl_2$ (1 m.mole) in aqueous ethanol (1:1:10ml) in the stoichiometric ratio [2Val:Zn:bipy] (Fig. 3). The solution was stirred constantly at room temperature for 4 hours then allowed to stand overnight. The

crystallized product was filtrated off and washed with aqueous ethanol.

2.3 X-ray crystallography (XRD)

The crystallographic features of the synthesized complexes were studied by X-ray diffraction (XRD) analysis *via* a GBC EMMA diffractometer with CuKα radiation (λ=0.154nm). The diffraction angle value (2θ) in the range 20°-70° was scanned at 2°/min with a step size of 0.02°. The analysis was carried out with beam acceleration at operating voltage and current of 35 kV and 28 mA, respectively.

2.4 UV–visible spectra

The absorption spectra of the synthesized complexes were recorded within the wavelength range of 250–850nm using a UV-visible PerkinElmer spectrometer. A halogen lamp was used as a light source coupled with a diffraction grating and photodiode detector. Light intensity calibration was performed by recording a baseline spectrum for a quartz tube filled in water. This calibration process eases the suppression of residual noise.

2.5 Fourier-transform infrared spectroscopy (FTIR)

The infrared reflection spectra of the synthesized complexes were obtained using a "reflected off" type of Perkin Elmer Spectrum 100 FTIR spectrometer in a wavenumber range of 400-4000 cm^{-1}. The samples were placed on a diamond crystal surface area and a pressure the arm was positioned and locked at a force of 100N in order to confirm the sample was touching evenly onto the crystal surface. Background correction was made before the collection of each spectrum.

2.6 Bioactivity assay

Calcium and zinc complexes were prepared and Ribeiro's modified medium (RMM) [37] with 0.35 mM phosphate and the Zn and Ca chelate were used for liquid media. The pH of the medium was adjusted to 6.4 (using KOH) and autoclaved before addition of phosphite, Zn chelate or Ca chelate. Media were dispensed into 9 cm diameter Petri-dishes (25 ml of liquid Ribeiro‟s modified medium).

The *P. cinnamomi* isolate MP94-48 was used in this study. Briefly, the isolate was grown on V8 agar plates for 7 days in the dark at 25°C, after which 5 mm diameter plugs were transferred to the RMM plates containing the different phosphite, Zn chelate and Ca chelate treatments. The *P. cinnamomi* isolate was grown in liquid RMM at an initial pH of 6.4 at different concentrations (0, 0.005, 0.40, 0.8 and 0.16, g/l) of phosphite or the chelate complexes. Inoculated plates (3 replications) were sealed with Parafilm® and incubated in the dark at 25°C without shaking. Mycelial growth was assessed by measuring the mycelial biomass dry weight after 7 days. Briefly, the mycelia were lifted out of the liquid, blotted dry with filter paper and dried in an oven at 70°C for one day before weighing. The EC50 values were computed from plots of the percent inhibition at different phosphite and Zn chelate or Ca chelateconcentrations compared to growth in the RMM without treatments.

3. Results and Discussion

3.1 Synthesis of Metal Complexes

In aerobic conditions, the complexes were made through interacting the metal salts with the corresponding ligands using a ratio of (1 MCl_2:1 2,2-bipyridine:2 sodium valinate) **(**Fig 3). The mixed ligand metal complexes synthesis can be expressed in the following equations:

$2Val\ H + 2NaOH \rightarrow 2Val\text{-}\ Na+ + H_2O$

$2Val\text{-}Na^+ + bipy + MCl_2 \rightarrow [M(Val)2(bipy)] + 4H_2O + NaCl$

M= Ca (II), Zn (II)

where bipy is 2,2-bipyridine and Val H is amino acid L-valine

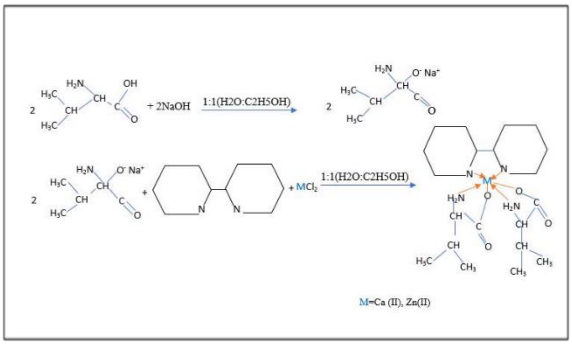


Fig. (3) Schematic representation of the preparation of the complexes [M(Val)2(bipy)]

The physicochemical properties, formulae weights and the melting points of the chemicals are listed in table (3). Both complexes were non-hygroscopic, stable at room temperature and white in color for both solid and liquid states. The results observed in this investigation suggest that the ligands acid L-valine and 2,2-bipyridine coordinate with either Ca(II) or Zn(II) form octahedral geometry. However, the yield of Zn complex was at 50% of basic materials. Whereas, the yield of Ca chelate was at 81% of the basic materials and is considered a good yield rate.

Table (3) Physicochemical properties of Ca and Zn complexes

Compounds	Molecular Weight	Physical state	% yield
Zn[(Val)2(bipy)]	453.80	White crystalline powder	81.39
Ca[(Val)2(bipy)]	428.50	White powder crystalline	55.55

3.2 Structural Characterization of Metal Complexes *XRD Analysis*

The crystalline structures of the Zn and Ca chelate complexes were determined by X-ray diffraction (XRD) (Fig. 4). The XRD results reported reflection faces at 77, 72, 89, 123, 143, 75, 100 and 99 nm for the Ca-complex, and 73, 75, 108, 80, 130, 103, 86 and 96 nm for the Zn-complex.

Both samples had the same patterns of structure indicating crystal structure symmetry. For example, the main peak of the Ca-chelate was located at 2θ=32.86°, which is very close to that of the Zn-chelate located at 2θ=32.9°. The principal diffraction peaks of both the Ca- chelate and Zn-chelate were placed in a wide range of temperatures between 25-50 °C. The intense and sharp peaks indicate the improved crystalline quality of the complexes.

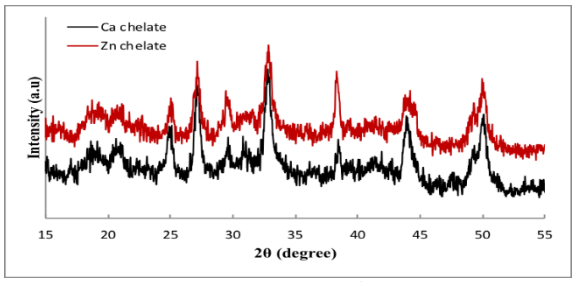


Fig. (4) X-ray crystallography data of the calcium and zinc chelate complexes

3.3 UV-visible spectra (UV-Vis)

Recorded absorption spectra of the calcium and zinc chelate complexes can be utilized to confirm their structure and to provide evidence that the electronic transitions occur. In both mixed ligand complexes (Zn complex and Ca complex), the absorption spectra displayed an absorption band around 280nm (Fig. 5) which could be attributed to (n→π*) transition. Also, both complexes did not show any d-d transitions because of their weakness.

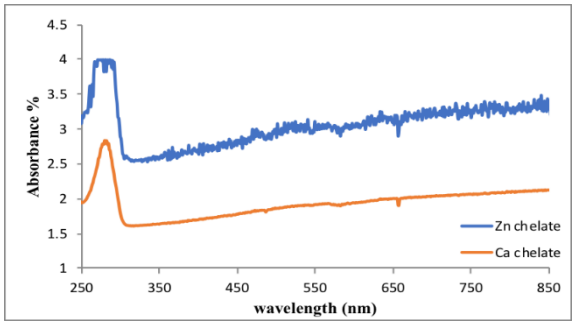
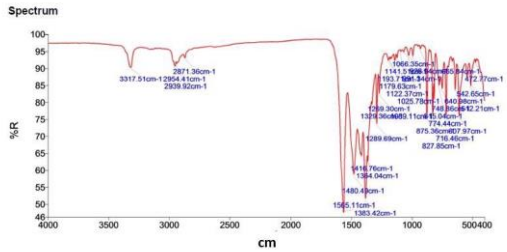


Fig. (5) Electronic spectral data of the Zn and Ca chelates

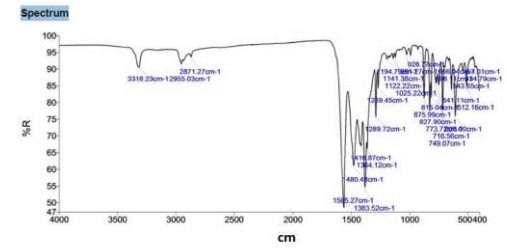
3.4 FTIR spectroscopy

The Zn and Ca complexes were also examined by FTIR analysis. This technique can predict the coordination of the complexes to their corresponding ligands. The most important IR band (peak) for both the Zn and Ca chelates was observed at a wavelength of 1565 cm^{-1} (Fig. 6a and b)

which reflects a vibrational mode for the υ (C=N) group of 2,2-bipyridine. This vibrational mode suggests that 2,2-bipyridine is similar to 1,10-phenanthroline, and is coordinated to the metal centers [38,39]. Another strong band for both chelates was observed around 3317 cm^{-1} and corresponds to the vibrational mode of (N-H) of the amine group. Two bands at 1480 cm^{-1} and 1364 cm^{-1} were related to υ (OCO) symmetry, which indicates the coordination of the carboxylic group to the central metal ion [40,41]. Moreover, other bands with low intensities in the spectra of both complexes were observed in the ranges of 610-542 and 472-401 cm^{-1} are due to metal-nitrogen υ (M-N) and metal–oxygen υ (M-O) stretching vibrations, respectively [42-44].



(a)



(b)

Fig. (6) Fourier-transform infrared spectroscopy data of (a) Ca chelate, and (b) Zn complex

3.4 Bioactivity assay

At 0.005 g/l of all chemical treatments the mycelial growth did not differ statistically from the control. For the Ca complex treatments at 0.04, 0.08, and 0.16 g/l as well as phosphite at 0.16 and 0.08 mycelial growth differed statistically to the control. In contrast, mycelial growth in all concentrations of Zn complex did not differ to the control. The Ca chelate reduced mycelial growth of *P. cinnamomi* more than any of the other treatments tested (table 4 and Fig. 7).

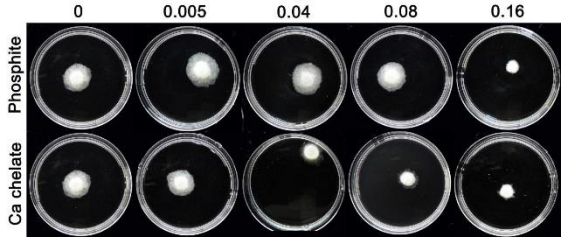


Figure 7.Colony growth of Phytophthora cinnamomi grown in liquid Ribeiro's modified medium in the presence of Ca chelate with Phosphite, the concentration of the chemical of g/l

Ligand complexes represent a novel group of candidates with potential as fungicides to control a range of plant pathogens including *Phytophthora* species. The XRD results showed similar patterns for both complexes indicating a similar crystal structure for both of them. This finding is consistent with that of [45] who indicated that similar complex patterns lead to similar crystal structures. The FT-IR results confirmed the presence of carboxylic and amine groups as well as metal-nitrogen υ(M-N) and metal–oxygen υ(M-O) in the metal complexes [46].Whilst, the UV-Vis results confirmed the presence of an aromatic ring. This result agrees with the findings of [47]. The yield of the Zn complex was at 50% of basic materials whilst the yield of the Ca chelate was better at 81% of basic materials. The higher yield of Ca complex makes it more suitable than the Zn complex as a potential fungicide. These results were consistent with numerous studies [48,49]. Good biological activity of mixed ligand complexes against pathogenic microorganisms in both animals and plants. The results of the present study show that Ca complex effectively inhibit the growth of *Phytophthora cinnamomi.* A possible explanation for the significant effect of Ca chelate might be that calcium ions stimulate may inhibit the growth of *Phytophthora*

species by suppression of sporangia formation [50,51].

References

[1] E. Canpolat, M. Kaya and S. Gür, "Synthesis, characterization of some Co(III) complexes with vicdioxime ligands and their antimicrobial properties", Turkish J. Chem., 28 (2004) 235-242.
[2] H.B. Gdr and N. Sharon, "Nomenclature and symbolism for aminoacids and peptides", Pure Appl. Chem., 56 (1984) 595-624.
[3] Aliyu and Isyaku, 2010, Shimazaki *et al.*, 2009).
[4] C. Kaes et al., "Bipyridine: the most widely used ligand: A review of molecules comprising at least two 2,2,,-bipyridine units", Chem. Rev., 100 (2000) 3553-3590.
[5] D.E. Fenton, "Biocoordination chemistry", Oxford University Press (UK, 2001).
[6] M. Girgaonkar and S. Shirodkar, "Synthesis, characterization and biological studies of Cu(II) and Ni (II) complexes with new bidentate Schiff's base ligands as 4-hydroxy-3-(1-(arylimino)ethyl)(arylimino)ethyl) Chromen-2-one", Res. J. Recent Sci., 1 (2012) 110-116.
[7] M. Shebl et al., "Binary and ternary copper(II) complexes of a new Schiff base ligand derived from 4-acetyl-5, 6-diphenyl-3(2H)- pyridazin one: Synthesis, spectral, thermal, antimicrobial and antitumor studies", J. Mol. Struct.,1145 (2017) 329-338.
[8] Dr. R.D. MADAN, "Modern Inorganic Chemistry," RAM NAGAR, NEW DELHI, 2000.
[9] J.C. Blair, E.M. Jones, in: H.S. Booth (Ed.), Inorganic Synthesis, vol. 1, McGraw-Hill Book Co., New York, 1939, p.35
[10] H.E. Spencer, M.W. Schmidt, J. Phys. Chem. 75 (1971) 2986.
[11] D. Armenatano, G. De Munno, F. Lioret, M. Julve, Cryst. Eng. Commun. 7 (2005) 57. 5- Tiexiang Fu, Jun Tao, Sens. Actuat. B 129 (2008) 339.
[12] K. Gao, J. Pearce, J. Jones, C. Taylor, *Environ. Geochem. Health* 21 (1999) 13 (Canpolat et al, 2004) Duncan, White, 2012, Jomová, Valko, 2011.
[13] J. R. J. Sorenson, Prog. Med. Chem. 26, 437 (1989) and refs. Therein
[14] K. D. Karlin, Z. Tyek´ar, Bioinorganic Chemistry of Copper, Chapman and Hall, New York (1993).
[15] T.P. N.Chaudhuri,Sorel, TetrahedronMulti-copper45, Oxidases,3 (1989). World Scientific, Singapore (1997).
[16] Koca, E.; Buyukasik, Y.; Cetiner, D.; Yilmaz, R.; Sayinalp, N.; Yasavul, U.; Uner, A. Copper deficiency with increased hematogones mimicking refractory anemia with excess blasts. *Leuk. Res.* 2008, *32*, 495-499.
[17] Twomey, P.J.; Reynolds, T.M.; Wierzbicki, A.S.; Viljoen, A. The relationship between serum copper and ceruloplasmin in routine clinical practice. *Int. J. Clin. Pract.* 2008, *62*, 485- 487.
[18] Prachayasittikul, S.; Suksrichavalit, T.; Isarankura-Na-Ayudhya, C.; Ruchirawat, S.; Prachayasittikul, V. Antimicrobial and antioxidative activities of 1-adamantylthio derivatives of 3-substituted pyridines. Excli J. 2008, 7, 63-70.
[19] ChemBio 2D, 3D Ultra, version 13.0, ChemBio Office (2012), Cambridge Soft Corporation, Cambridge, MA, USA, 2012
[20] Anhalt JP, Washington JA. In Laboratory procedure in clinical Microbiology, 2nd ed., Washington JA (ed.). Springer: New York, 1985; 281-313.
[21] Bull. Chem. Soc. Ethiop. 2010, 24(3).
[22] J.C. Blair, E.M. Jones, in: H.S. Booth (Ed.), Inorganic Synthesis, vol. 1, McGraw-Hill Book Co., New York, 1939, p. 35, in:.
[23] H.E. Spencer, M.W. Schmidt, J. Phys. Chem. 75 (1971) 2986.
[24] D. Armenatano, G. De Munno, F. Lioret, M. Julve, Cryst. Eng. Commun. 7 (2005) 57.
[25] Tiexiang Fu, Jun Tao, Sens. Actuat. B 129 (2008) 339.
[26] C.J.Ballahausen, "Introduction to Ligand Field Theory", Mc Graw-Hill Book Co., New York (1962).
[27] Lever .A. B.P. "*Inorganic spectroscopy*" Elsevier publishing company, London, New York. (1984).
[28] Nishida. Y., Niinuma. A., Abe. K., Inorg. Chem. Commun. 12 198–200(2009).
[29] Shriver & Atkins*" Inorganic Chemistry*", 3rd Ed., , Freeman, (1999).
[30] J. Fujita A. E. Martell, and K. Nakamoto, J. Chem. Phys., 36, 324, 331 (1962) . 29- R. D. Hancock and D. A. Thornton, J. Mol. Struct., 6, 441 (1970).
[31] Nakamoto K. "Infrared and Raman Spectra of Inorganic and Coordination Compounds Applications in Coordination Organometallic and Bioinorganic Chemistry,4 th Ed.,244 (1997)
[32] A. M. Rahman, M. I. Choudhary and W. J. Thomsen, "Bio-assay Techniques for Drug Development, Harwood Academic, Netherlands, 2001.
[33] Cotton FA, Wilkinson G. and Murillo CA. Advanced Inorganic Chemistry. 1999;1355.
[34] Miessler, Gray L. and Tarr DA. coordination complexes inorganic chemistr.1999;642.

[35] Kaim, W.; Schwederski, B. Bioinorganic chemistry: inorganic elements of life, John Wiley and Sons: London;1996; pp. 39-262.
[36] Xiao-Ming, C.; Bao-Hui, Y.; Xiao, C.H.; Zhi-Tao, X. J. chem. Soc., Dalton Trans. 1996, 3465.
[37] Merck and Co.; Merck Index, 2006, Fourteen Edition, p. 1974.
[38] Usp 27-NF 22, Pharmacopeia, 2004, p. 393.
[39] J. R. Anacona and I. Radriguez, (synthesis and antibacterial activity of cephalexin metal complexes), J. Coord. Chem., 2004, 57(15), p.1263-1269.
[40] O, Grady F, Finch RG, Lambert HP and Greenwood D(1997). Antibiotic and chemotherapy, Anti-infective agents and their use in therapy, 7th edition Churchill Livingstone, New York, pp.202-15.
[41] Beta Hydroxy Acids in Cosmetics. FDA. March 7, 2000.
[42] Yates, R.L., and D.C. Havery, "Determination of Phenol, Resorcinol, Salicylic Acid and Alpha-Hydroxy Acids in Cosmetic Products and Salon Preparations," Journal of Cosmetic Science, vol. 50, pp. 315-325, 1999.
[43] Salicylic Acid Topical. Medline Plus. September 1, 2008.
[44] Balali-Mood M, Balali-Mood K. Salicylic acid. "Inter-Organization Programmed for the Sound Management of Chemicals" INCHEM. August 1996.
[45] Diarmind Jeffreys, "Aspirin the remarkable story of a wonder drug", By: Bloomsulary, pp.38-40, 2005.
[46] Philip A. Mackowiak, "Brief History of Antipyretic therapy clinical infection diseases", London, pp.154-156, 2000.
[47] Peter C. Ford, (2007), "Reaction Mechanisms of Inorganic and Organometallic Systems" 3rd Ed., Oxford University Press, p. 532.
[48] M.R. Mahmoud, A.M. Hamman and S.A. Ibrahim; (1984), Z.Phys. Chem., Leipzig, p. 265-302.
[49] Zahid H. C. and Maimoon F. J., (2000), "synthesis, characterization and biological evaluation of Co(ii), Cu(ii), Ni(ii) and Zn(ii) complexes with cephradine" Vol. 7, Nr.5, pp255-269.
[50] M.R. Mahmoud, A.M. Hamman and S.A. Ibrahim; (1984), Z.Phys. Chem., Leipzig, p. 265-302.
[51] K. R. Adam, M. Antolovich, D. S. Baldwin, P. A. Duckworth, A. J. Leong, L. F. Lindoy, M. McPartlin, P. A. Rasker, (1993), J. Chem. Soc., Dalton Trans. p:1013.

Table 1 the different properties of valine

Compound	Molecular Formula	Density (g/cm^3)	Molar Mass (g/mol)	Solubility	Acidity (pKa)	Physical state	Melting point (°C)
Valine	$C_5H_{11}NO_2$	1.316	117.15	Soluble in water	2.32	White crystalline powder	298
					(carboxyl)		
					9.62		
					(amino)		

Table 2 the different properties of 2,2′-bipyridine

Compound	Molecular formula	Density (g/cm^3)	Molar mass (g/mol)	Solubility	Acidity (pKa)	Physical state	Melting point °C
2,2'-bipyridine	$C_2H_8N_{10}$	1.316	156.18	Slightly Soluble in water	9.67	White crystalline powder	70-73

Table (4) Dry weights (±SE) and EC50 values of *Phytophthora cinnamomi* grown in liquid Ribeiro's modified medium For each column, values with the same letter are not significantly different (P≤ 0.05)

Treatments	Concentration (g/l)	Dry Weight (µg)	EC 50
Control	0		
Phosphite	0.005	2.14+0.21 d	100
Ca complex	0.04	2.14+0.08 d	100
Ca complex	0.08 0.16	1.90+0.12 d	87
Ca complex	0.005 0.04	1.57+0.21c 1.38+0.33b 2.09+0.53d 1.42+0.33b 1.28+0.14b	73 67 100 67
Ca complex, Zn	0.08 0.16	1.04+0.25a 2.14+0.08 d	60 47
complex	0.005	2.00+0.08 d	100
Zn complex	0.04	2.00+0.21 d	93
Zn complex	0.08	1.85+0.16 d	93
Zn complex	0.16		87

